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Study Guide Chemistry Stoichiometry Answer The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. ____ 2. Stoichiometry is based on the law of conservation of mass. ____ 3.

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Chemistry Holt Study Guide Mole Concept

The relative masses were obtained by multiplying the atomic ratios and atomic masses. You can see that a sample of N₂O weighing 44.02 grams contains 28.02 g of nitrogen and 16.00 g of oxygen. The mass percent of each element is calculated from its relative mass divided by the sum of the relative masses.

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Given a chemical reaction, stoichiometry tells us what quantity of each reactant we need in order to get enough of our desired product. Because of its real-life applications in chemical engineering as well as research, stoichiometry is one of the most important and fundamental topics in chemistry.

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